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Redox and electrochemistry finding the oxidation state worksheet answers

For an electrolytic cell $E < 0$, $\Delta G > 0$. Electrons flow from anode to cathode (this is always the case). For an electrolytic cell however, this flow is not spontaneous but must be driven by an external power source. In an electrolytic cell, the anode has the "+" sign. Page ID158277 No headers Was this article helpful? If you're seeing this message, it means we're having trouble loading external resources on our website. If you're behind a web filter, please make sure that the domains *.kastatic.org and *.kasandbox.org are unblocked. Course Handouts » Chemistry » Unit Fifteen - Redox and Electrochemistry » Classwork and Homework Handouts Classwork and Homework Handouts Classwork and Homework Handouts NEED HELP DOWNLOADING: doc file: You need the Microsoft Word program, a free Microsoft Word viewer, or a program that can import Word files in order to view this file. To learn more about the free Microsoft Word Viewer, visit the Microsoft Word website. docx file: You need the Microsoft Word program, the Microsoft Word app, or a program that can import Word files in order to view this file. To learn more about the free Microsoft Word app, visit the Microsoft store. Mr. Abraham's Site Penfield High School 25 High School Drive Penfield, NY 14526 andriano_cz / Getty Images Electrochemical reactions involve the transfer of electrons. Mass and charge are conserved when balancing these reactions, but you need to know which atoms are oxidized and which atoms are reduced during the reaction. Oxidation numbers are used to keep track of how many electrons are lost or gained by each atom. These oxidation numbers are assigned using the following rules. The convention is that the cation is written first in a formula, followed by the anion. For example, in NaH, the H is H⁻; in HCl, the H is H⁺. The oxidation number of a free element is always 0. The atoms in He and N₂, for example, have oxidation numbers of 0. The oxidation number of a monatomic ion equals the charge of the ion. For example, the oxidation number of Na⁺ is +1; the oxidation number of N³⁻ is -3. The usual oxidation number of hydrogen is +1. The oxidation number of hydrogen is -1 in compounds containing elements that are less electronegative than hydrogen, as in CaH₂. The oxidation number of oxygen in compounds is usually -2. Exceptions include OF₂ because F is more electronegative than O, and BaO₂, due to the structure of the peroxide ion, which is [O-O]²⁻. The oxidation number of a Group IA element in a compound is +1. The oxidation number of a Group IIA element in a compound is +2. The oxidation number of a Group VIIA element in a compound is -1, except when that element is combined with one having a higher electronegativity. The oxidation number of Cl is -1 in HCl, but the oxidation number of Cl is +1 in HOCl. The sum of the oxidation numbers of all of the atoms in a neutral compound is 0. The sum of the oxidation numbers in a polyatomic ion is equal to the charge of the ion. For example, the sum of the oxidation numbers for SO₄²⁻ is -2.

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